

# How to draw Lewis Dot Structures for all Covalents

- ❑ Step 1 - Count the total Valance elec.
- ❑ Step 2 - Draw a structure with only single bonds
- ❑ Step 3 - Count the Val. elec. needed
- ❑ Step 4 - Compare the Val. elect. you need & what you have
  - If they are the same - Hooray!
  - If you are off by 2 : put in a double bond
  - if you are off by 4 : put in a triple or 2 doubles bonds
- ❑ Step 5 - Remember Usually (but not always):
  - $(\# \text{ of Val. Elect}) + (\text{bonds}) = 8$
  - For example: Nitrogen has 5 val. elect. and often has 3 bonds