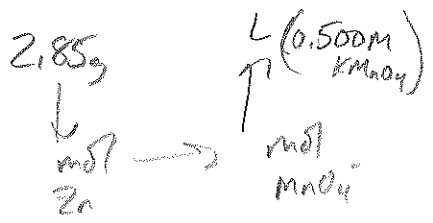
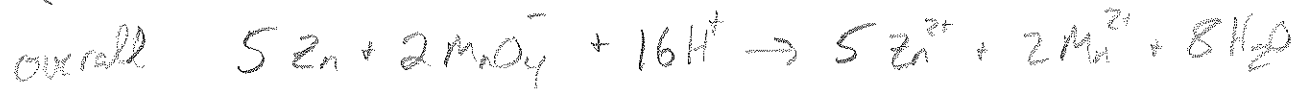
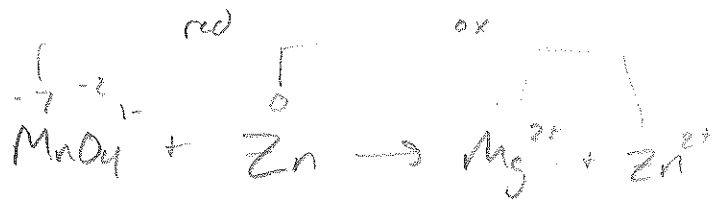


Ch 18  
# 103



$$2.85\text{g} \left( \frac{1\text{mol}}{65.4\text{g}} \right) \left( \frac{2\text{mol MnO}_4^-}{5\text{mol Zn}} \right) \left( \frac{1\text{L}}{0.500\text{mol}} \right) \left( \frac{1000\text{mL}}{1\text{L}} \right) = 34.9\text{mL}$$

$\uparrow$   
 0.500M  $\text{KMnO}_4$   
 solution